



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CHEMISTRY

0620/11

Paper 1 Multiple Choice (Core)

May/June 2018

45 minutes

Additional Materials: Multiple Choice Answer Sheet
 Soft clean eraser
 Soft pencil (type B or HB is recommended)

* 9 5 6 3 3 3 2 1 6 8 *

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

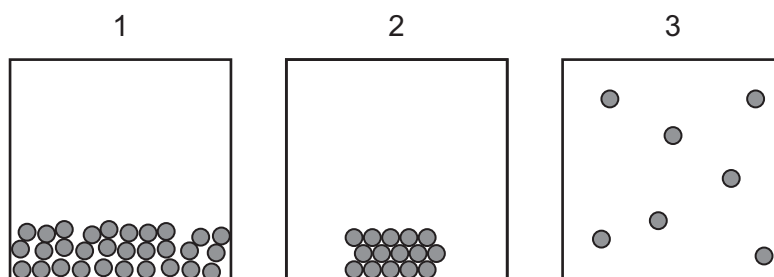
A copy of the Periodic Table is printed on page 16.

Electronic calculators may be used.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **14** printed pages and **2** blank pages.

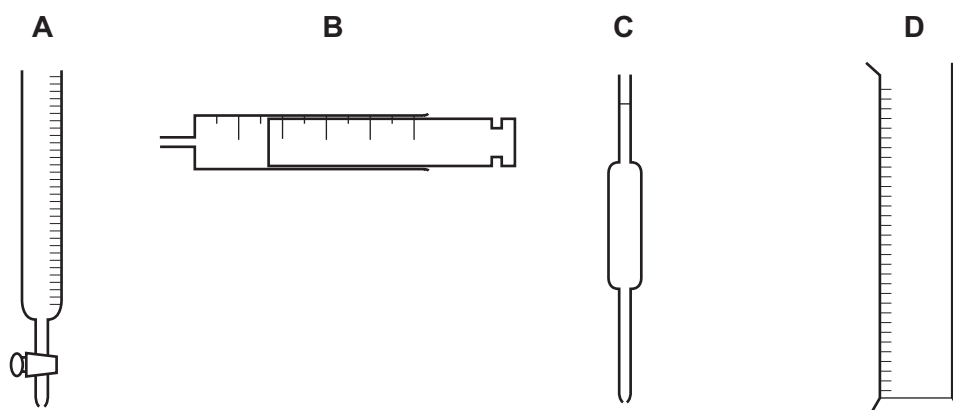
1 The diagrams show particles in a container.



Which two diagrams show the process of evaporation?

- A** 1 → 2 **B** 1 → 3 **C** 2 → 3 **D** 3 → 1

2 Which piece of apparatus is used to measure exactly 26.3 cm³ of a liquid?



3 The melting points and boiling points of pure substances W, X and Y are shown.

| | W | X | Y |
|------------------|------|-----|------|
| melting point/°C | -114 | 115 | -101 |
| boiling point/°C | 78 | 445 | -34 |

The substances are chlorine, ethanol and sulfur.

Which row identifies W, X and Y?

| | W | X | Y |
|----------|----------|----------|----------|
| A | chlorine | ethanol | sulfur |
| B | ethanol | sulfur | chlorine |
| C | sulfur | chlorine | ethanol |
| D | sulfur | ethanol | chlorine |

4 In which atom is the number of protons equal to the number of neutrons?

- A ^{40}Ar B ^{19}F C ^{23}Na D ^{16}O

5 Which row identifies an alloy, a pure metal and a non-metal?

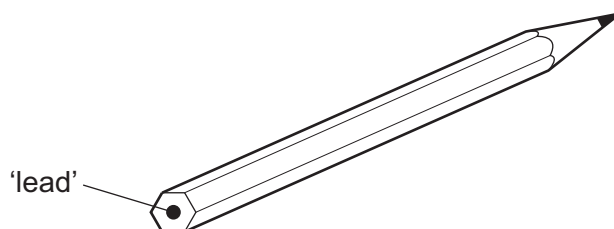
| | alloy | pure metal | non-metal |
|----------|--------|------------|-----------|
| A | brass | carbon | copper |
| B | brass | copper | carbon |
| C | copper | brass | carbon |
| D | copper | carbon | brass |

6 A covalent molecule Q contains exactly six shared electrons.

What is Q?

- A ammonia, NH_3
 B chlorine, Cl_2
 C methane, CH_4
 D water, H_2O

7 The 'lead' in a pencil is made of a mixture of graphite and clay.



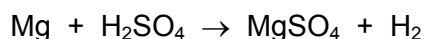
When the percentage of graphite is increased, the pencil slides across the paper more easily.

Which statement explains this observation?

- A Graphite has a high melting point.
 B Graphite is a form of carbon.
 C Graphite is a lubricant.
 D Graphite is a non-metal.

- 8 The equation for the reaction between magnesium and dilute sulfuric acid is shown.

The M_r of $MgSO_4$ is 120.



Which mass of magnesium sulfate is formed when 12g of magnesium completely reacts with dilute sulfuric acid?

- A** 5g **B** 10g **C** 60g **D** 120g
- 9 What is observed at each electrode when molten lead(II) bromide is electrolysed using platinum electrodes?

| | negative electrode | positive electrode |
|----------|-----------------------------|-----------------------------|
| A | bubbles of a colourless gas | bubbles of a brown gas |
| B | bubbles of a colourless gas | bubbles of a colourless gas |
| C | shiny grey liquid | bubbles of a brown gas |
| D | shiny grey liquid | bubbles of a colourless gas |

- 10 Which gas is used as a fuel?

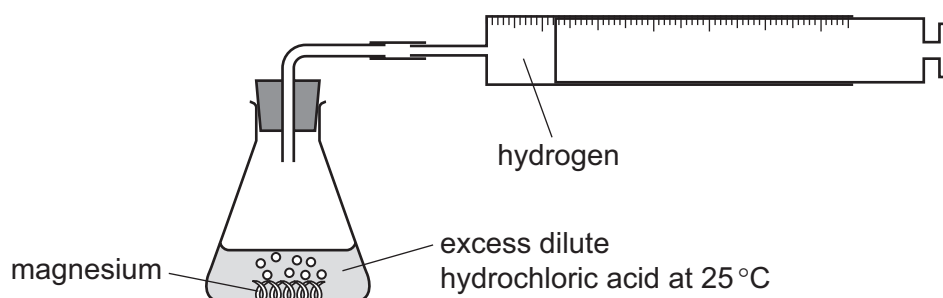
- A** argon
B hydrogen
C nitrogen
D oxygen

- 11 Burning fuels is an exothermic reaction.

What is meant by the term *exothermic*?

- A** A gas is produced.
B Energy is released.
C Heat is absorbed.
D The mass of the fuel decreases.

12 The diagram shows a rate of reaction experiment.



Increasing the concentration of the acid and increasing the temperature both affect the rate of reaction.

Which row is correct?

| | increase the concentration of acid | increase the temperature |
|----------|------------------------------------|---------------------------|
| A | decrease rate of reaction | decrease rate of reaction |
| B | decrease rate of reaction | increase rate of reaction |
| C | increase rate of reaction | decrease rate of reaction |
| D | increase rate of reaction | increase rate of reaction |

13 Water is added to anhydrous copper(II) sulfate.

What happens during the reaction?

- A** The copper(II) sulfate turns blue and the solution formed gets colder.
- B** The copper(II) sulfate turns blue and the solution formed gets hotter.
- C** The copper(II) sulfate turns white and the solution formed gets colder.
- D** The copper(II) sulfate turns white and the solution formed gets hotter.

14 Which equation shows an oxidation reaction?

- A** $C + O_2 \rightarrow CO_2$
- B** $CaCO_3 \rightarrow CaO + CO_2$
- C** $CaO + 2HCl \rightarrow CaCl_2 + H_2O$
- D** $N_2O_4 \rightarrow 2NO_2$

15 Dilute nitric acid is added to a solid, F.

A gas, G, is produced which is denser than air and extinguishes a burning splint.

What are F and G?

| | solid F | gas G |
|----------|-------------------|----------------|
| A | calcium | hydrogen |
| B | calcium carbonate | carbon dioxide |
| C | calcium hydroxide | hydrogen |
| D | calcium oxide | carbon dioxide |

16 Which statement about oxides is correct?

- A** A solution of magnesium oxide has a pH less than pH 7.
- B** A solution of sulfur dioxide has a pH greater than pH 7.
- C** Magnesium oxide reacts with nitric acid to make a salt.
- D** Sulfur dioxide reacts with hydrochloric acid to make a salt.

17 Which methods are suitable for preparing **both** zinc sulfate and copper(II) sulfate?

- 1 reacting the metal oxide with warm dilute aqueous sulfuric acid
- 2 reacting the metal with dilute aqueous sulfuric acid
- 3 reacting the metal carbonate with dilute aqueous sulfuric acid

- A** 1, 2 and 3 **B** 1 and 2 only **C** 1 and 3 only **D** 2 and 3 only

18 Two salt solutions, X and Y, are tested.

The table shows the results.

| test | X | Y |
|---|--------------------------|------------------------------|
| a few drops of aqueous sodium hydroxide are added | green precipitate formed | red-brown precipitate formed |
| a few drops of dilute nitric acid and a few drops of barium nitrate are added | no change seen | white precipitate formed |
| a few drops of dilute nitric acid and a few drops of silver nitrate are added | white precipitate formed | no change seen |

What are X and Y?

| | X | Y |
|----------|--------------------|--------------------|
| A | iron(II) chloride | iron(III) sulfate |
| B | iron(III) chloride | iron(III) sulfate |
| C | iron(II) sulfate | iron(III) chloride |
| D | iron(III) sulfate | iron(III) chloride |

19 Which element is in the same period of the Periodic Table as silicon?

- A** germanium
- B** scandium
- C** sodium
- D** strontium

20 Which statement about the halogens is correct?

- A** A sample of bromine reacts with potassium chloride solution.
- B** A sample of bromine reacts with potassium iodide solution.
- C** A sample of chlorine has a higher density than a sample of bromine.
- D** A sample of chlorine is a darker colour than a sample of bromine.

21 Which row shows the catalytic activity of transition elements and their compounds?

| | catalytic activity of transition elements | catalytic activity of compounds of transition elements |
|----------|---|--|
| A | good | good |
| B | good | poor |
| C | poor | good |
| D | poor | poor |

22 Which statement about the noble gases is **not** correct?

- A** Noble gases are diatomic molecules.
- B** Noble gases are unreactive gases.
- C** Noble gases have full outer electron shells.
- D** The noble gas argon is used in lamps.

23 The following statements are made about the metals copper, iron, magnesium and zinc.

- 1 Their oxides are acidic.
- 2 They all conduct electricity in the solid state.
- 3 They all have high melting points.
- 4 They all react with dilute acids to form hydrogen.

Which statements are correct?

- A** 1 and 2 **B** 1 and 4 **C** 2 and 3 **D** 3 and 4

24 Three metals, X, Y and Z, were reacted with water.

The oxides of the same three metals were also heated strongly with carbon.

The results are shown.

| metal | reaction of the metal with water | reaction of the metal oxide with carbon |
|-------|---|---|
| X | vigorous reaction with cold water | no reaction |
| Y | no reaction | metal and carbon dioxide produced |
| Z | no reaction observed with cold water but reaction observed with steam | no reaction |

What is a correct conclusion about X, Y and Z?

- A X is sodium and Y is magnesium.
- B X is the least reactive and Y is the most reactive.
- C Z is less reactive than Y.
- D Z is magnesium and Y is copper.

25 In a blast furnace, iron ore is mixed with coke and limestone, and heated in hot air.

Compound R is formed. Compound R then reduces the iron ore to iron.

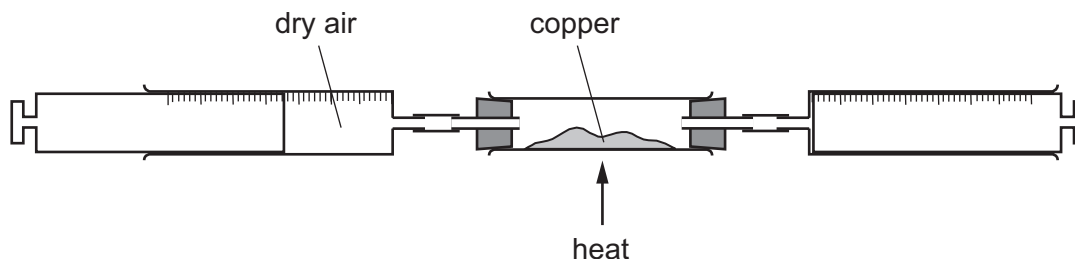
Which equation shows the formation of compound R?

- A $C + O_2 \rightarrow CO_2$
- B $CO_2 + C \rightarrow 2CO$
- C $CaCO_3 \rightarrow CaO + CO_2$
- D $CaO + SiO_2 \rightarrow CaSiO_3$

26 Which statement explains why aluminium is used in the manufacture of aircraft?

- A It conducts heat well.
- B It has a low density.
- C It is a good conductor of electricity.
- D It is easy to recycle.

- 27 Dry air is passed over hot copper until all the oxygen has reacted.



The volume of gas at the end of the reaction is 120 cm^3 .

What is the starting volume of dry air?

- A** 132 cm^3 **B** 152 cm^3 **C** 180 cm^3 **D** 570 cm^3
- 28 A steel bicycle which had been left outdoors for several months was starting to rust.
- What would **not** reduce the rate of corrosion?
- A** Remove the rust and paint the bicycle.
B Remove the rust and store the bicycle in a dry shed.
C Remove the rust and wipe the bicycle with a clean, damp cloth.
D Remove the rust and wipe the bicycle with an oily cloth.
- 29 Which statements about water are correct?
- 1 Household water contains dissolved salts.
 - 2 Water for household use is filtered to remove soluble impurities.
 - 3 Water is treated with chlorine to kill bacteria.
 - 4 Water is used in industry for cooling.
- A** 1, 2, 3 and 4
B 1, 2 and 3 only
C 1, 3 and 4 only
D 2, 3 and 4 only
- 30 Farmers use fertilisers to replace minerals in the soil that have been removed by the crops they grow.

Which elements in the soil are replaced by adding fertilisers?

- A** Ca, P, O **B** K, O, S **C** N, K, P **D** N, O, S

31 Which statement is correct?

- A Atmospheric carbon dioxide is not a cause of climate change.
- B Atmospheric methane is produced by respiration.
- C Burning natural gas decreases the level of carbon dioxide in the atmosphere.
- D Decomposition of vegetation causes an increase in atmospheric methane.

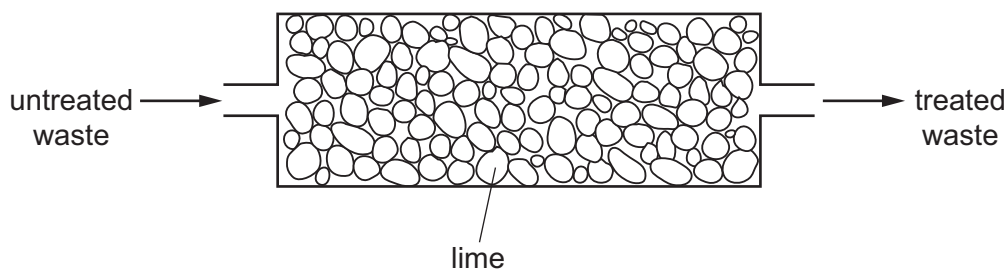
32 Which statement about sulfur and its compounds is **not** correct?

- A Sulfur dioxide is used as a food preservative.
- B Sulfur dioxide turns acidified aqueous potassium manganate(VII) from purple to colourless.
- C Sulfur forms a basic oxide.
- D Sulfur is used in the manufacture of sulfuric acid.

33 Which process is used to convert limestone (calcium carbonate) into lime?

- A electrolysis
- B fractional distillation
- C incomplete combustion
- D thermal decomposition

34 Lime is used to treat an industrial waste.



Which change occurs in the treatment?

| | untreated waste | | treated waste |
|----------|-----------------|---|---------------|
| A | acidic | → | neutral |
| B | alkaline | → | acidic |
| C | alkaline | → | neutral |
| D | neutral | → | acidic |

35 What is **not** the correct use of the fraction named?

| | name of fraction | use |
|----------|------------------|------------------------|
| A | fuel oil | making waxes |
| B | gas oil | fuel in diesel engines |
| C | kerosene | jet fuel |
| D | naphtha | making chemicals |

36 Four organic compounds are listed.

ethane

ethanoic acid

ethanol

ethene

Which bond do all four compounds contain?

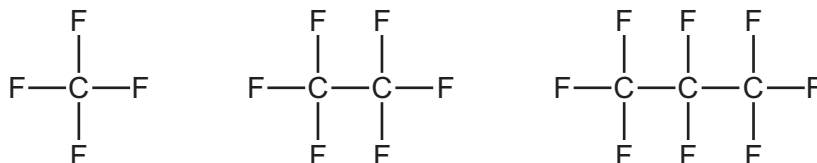
A C–C

B C–H

C C–O

D O–H

37 The first three members of a homologous series are shown.



Why do these molecules represent a homologous series?

A because they contain fluorine and carbon atoms

B because they have saturated bonds

C because they have the same functional group

D because they react differently from each other

38 Which substances can be obtained by cracking hydrocarbons?

A ethanol and ethene

B ethanol and hydrogen

C ethene and hydrogen

D ethene and poly(ethene)

39 Which reaction is used to make ethanol?

- A adding steam to ethene
- B addition polymerisation
- C fractional distillation of petroleum
- D reacting ethene with aqueous bromine

40 Polymers are long-chain molecules made from small molecules linked together.

Four polymers or types of polymer are listed.

- 1 carbohydrates
- 2 nylon
- 3 proteins
- 4 *Terylene*

Which of these polymers or types of polymer are synthetic?

- A 1 and 3 B 1 and 4 C 2 and 3 D 2 and 4

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The Periodic Table of Elements

| | | Group | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
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| I | II | III | IV | V | VI | VII | VIII | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| 3 Li lithium 7 | 4 Be beryllium 9 | 11 Na sodium 23 | 12 Mg magnesium 24 | 19 K potassium 39 | 20 Ca calcium 40 | 37 Rb rubidium 85 | 55 Cs caesium 133 | 87 Fr francium — | 1 H hydrogen 1 | 2 He helium 4 | 5 B boron 11 | 6 C carbon 12 | 7 N nitrogen 14 | 8 O oxygen 16 | 9 F fluorine 19 | 10 Ne neon 20 | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| 11 Na sodium 23 | 12 Mg magnesium 24 | 13 Al aluminium 27 | 14 Si silicon 28 | 15 P phosphorus 31 | 16 S sulfur 32 | 17 Cl chlorine 35.5 | 18 Ar argon 40 | 19 K potassium 39 | 20 Ca calcium 40 | 21 Sc scandium 45 | 22 Ti titanium 48 | 23 V vanadium 51 | 24 Cr chromium 52 | 25 Mn manganese 55 | 26 Fe iron 56 | 27 Co cobalt 59 | 28 Ni nickel 59 | 29 Cu copper 64 | 30 Zn zinc 65 | 31 Ga gallium 70 | 32 Ge germanium 73 | 33 As arsenic 75 | 34 Se selenium 79 | 35 Br bromine 80 | 36 Kr krypton 84 | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
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